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MAGEMin, an Efficient Gibbs Energy Minimizer: Application to Igneous Systems

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Key Points:

- A new, parallel, Gibbs energy minimization approach is presented to compute multiphase multicomponent equilibria
- It predicts parameters like stable phases, melt content, or seismic velocities as a function of chemistry and temperature/pressure conditions
- Examples and benchmark cases are presented that apply the approach to magmatic systems

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Abstract Prediction of stable mineral equilibria in the Earth's lithosphere is critical to unravel the tectonomagmatic history of exposed geological sections. While the recent advances in geodynamic modeling allow us to explore the dynamics of magmatic transfer in solid mediums, there is to date no available thermodynamic package that can easily be linked and efficiently be accounted for the computation of phase equilibrium in magmatic systems. Moreover, none of the existing tools fully exploit single point calculation parallelization, which strongly hinders their applicability for direct geodynamic coupling or for thermodynamic database inversions. Here, we present a new Mineral Assemblage Gibbs Energy Minimizer (MAGEMin). The package is written as a parallel C library, provides a direct Julia interface, and is callable from any petrological/geodynamic tool. For a given set of pressure, temperature, and bulk-rock composition MAGEMin uses a combination of linear programming, extended Partitioning Gibbs Energy and gradient-based local minimization to compute the stable mineral assemblage. We apply our new minimization package to the igneous thermodynamic data set of Holland et al. (2018), https://doi.org/10.1093/petrology/egy048 and produce several phase diagrams at supra-solidus conditions. The phase diagrams are then directly benchmarked against THERMOCALC and exhibit very good agreement. The high scalability of MAGEMin on parallel computing facilities opens new horizons, for example, for modeling reactive magma flow, for thermodynamic data set inversion, and for petrological/geophysical applications.

Plain Language Summary Understanding magmatic systems requires knowing how rocks melt. Because a single melting experiment can easily take weeks, it is impossible to do enough experiments to cover the whole range of pressure, temperature, and composition relevant for magmatic systems. We therefore need a way to interpolate in between conditions that are not directly covered by the experiments. It is long known that the best way to perform such interpolation is by using basic thermodynamic principles. For magmatic systems, this requires a well-calibrated thermodynamic melting model. It also requires an efficient computational tool to predict the most stable configuration of minerals and melt. Since the 1980s, a number of such computational tools have been developed to perform a so-called Gibbs energy minimization. These tools work very well for simpler systems but become very slow for recently developed, more realistic, melting models. Here, we describe a new method that combines some ideas of the previous methods with a new algorithm. Our method is faster and takes advantage of modern computer architectures. It can predict rock properties such as densities, seismic velocities, melt content, and chemistry. It can therefore be used to link physical observations with hard rock data of magmatic systems.

1. Introduction

The thermodynamic modeling of equilibrium mineral assemblages is a crucial tool for understanding the solid Earth. Mineral equilibrium modeling can be used in an inverse sense to make inferences about magmatic and tectonic processes based on the rocks that they generated. Used in a forward-modeling sense, our capacity to simulate Earth processes is greater if we can model the most stable mineral assemblage under given conditions, since the mineral assemblage controls or contributes to the thermodynamic, chemical, and rheological properties of the rock package. Such modeling thus forms a key step in linking geophysical observations with petrological constraints and to assess the effect of mineral reactions on deformation of the lithosphere. Even when geological systems are not always at equilibrium, nonequilibrium effects tends to move the system toward equilibrium

(Lasaga, 1986) and as such it remains crucial to be able to efficiently model the equilibrium state (e.g., Hou et al., 2021).

To model mineral equilibria, an equation of state is needed for each mineral or fluid phase that might potentially be stable under the conditions of interest. The equation of state describes the calorimetric and volumetric properties of the phase as a function of its pressure, temperature, composition, and state of order. A phase may be considered to have anything from one compositional component (a pure phase) to the maximum number of components in which the rock system is to be modeled. It may or may not contain dimensions of order-disorder. There are several different thermodynamic data sets currently in use that comprise collections of such equations of state, usually aimed at modeling a subset of terrestrial mineral equilibria; for example, those of R. W. White et al. (2014) (equilibria in metapelites), Green et al. (2016) (equilibria in metabasites), and Stixrude and Lithgow-Bertelloni (2011, 2021) (equilibria among mantle phases). Each collection is calibrated with some degree of internal consistency. In this contribution, we use a version of the thermodynamic data set of Holland et al. (2018), which incorporates the internally consistent data set of end-member thermodynamic properties of Holland and Powell (2011) and Tomlinson and Holland (2021). However, our method can also be applied to any other thermodynamic data sets.

Mineral equilibrium calculations for geological applications commonly assume that pressure and temperature are the independent variables in the problem, rather than their conjugates, volume, and entropy. Therefore, the equilibrium compositions and states of order for the phases in a model equilibrium are found by minimizing the Gibbs energy, G, of the assemblage. However, minimization of the Gibbs energy in multicomponent and multiphase systems remains one of the most challenging global optimization problems, not only in the fields of metamorphic petrology (e.g., Lanari & Duesterhoeft, 2018) but also for chemical engineering (e.g., Fateen, 2016; Lothenbach et al., 2019) and for the nuclear industry (e.g., Piro, 2011; Piro, Banfield, et al., 2013; Piro, Simunovic, et al., 2013). Because the problem is intrinsically multidimensional, nonlinear and nonconvex, minimization strategies are not guarantee to obtain the global minimum of the Gibbs energy of the system. Consequently, numerous Gibbs energy minimization strategies are used depending on the problem dimensionality (number of chemical components) and complexity of the equations of state. This includes, but is not limited to, equality- and nonequality-constrained linear least squares (e.g., Ghiorso, 1983, 1985), linear programming and nonlinear optimization methods (e.g., de Capitani & Brown, 1987), discretization of equations of state in composition-order space combined with linear programming (e.g., Connolly, 1990, 2005), linear programming and Partitioning Gibbs Energy (PGE; e.g., Kruskopf & Visuri, 2017; Piro, 2011; Piro, Banfield, et al., 2013; Piro, Simunovic, et al., 2013), metaheuristic optimization methods (e.g., Burgos-Solórzano et al., 2004; Çetin & Keçebaş, 2021; Teh & Rangaiah, 2002), and Lagrangian formulations (e.g., Piro & Simunovic, 2016; W. White et al., 1958).

In the geosciences, a number of petrological tools have been developed to predict phase equilibria, study phase relations, and produce phase diagrams, for example, Gibbs (Spear, 1988), THERMOCALC (Powell & Holland, 1988), Perple_X (Connolly, 1990, 2005), Theriak-Domino (de Capitani & Brown, 1987; de Capitani & Petrakakis, 2010), MELTS and pMELTS (Asimow & Ghiorso, 1998; Ghiorso, 1983, 1985; Ghiorso & Sack, 1995), and GeoPS (Xiang & Connolly, 2021). In general, they fall into two categories (Connolly, 2017; Lanari & Duesterhoeft, 2018): phase equilibrium calculators and Gibbs energy minimizers.

The first category (e.g., THERMOCALC and Gibbs) equates the chemical potentials of components in a specified set of phases, to calculate what compositions and states of order the phases must have to be in equilibrium with each other under the specified conditions. The user may investigate any set of phases for which equations of state are present within the thermodynamic data set. Conditions to be specified might include pressure, temperature, bulk system composition, or partial phase compositions. Univariant reactions or other phase field boundaries are calculated using geometric constraints (Schreinemakers analysis or related rules) combined with experience and a priori knowledge of the petrological system. This approach allows a wide range of phase diagrams to be calculated and facilitates the user in exploring any equilibrium of interest, whether stable or metastable. However, in many applications in the geosciences, the only equilibrium of interest is that of the most stable equilibrium at given bulk system composition—or, in geological terms, the most stable equilibrated mineral assemblage in a given bulk-rock composition. The phase equilibrium calculator approach is not optimal for this purpose, since it depends on the user anticipating all of the phases that might appear in the stable

equilibrium. In complex systems, even an expert user may easily overlook the presence of a phase in a given region of the diagram and consequently mistake a metastable assemblage for the stable one. Programs in the second category (e.g., MELTS, pMELTS, Theriak-Domino, Perple_X, and GeoPS) are designed specifically to predict the most stable assemblage in a given bulk-rock composition. At each point on a pressure-temperature grid, these programs explore all possible equilibria among subsets of the phases in a large prespecified list, potentially including all the phases represented in the thermodynamic data set. They return the subset of these phases that yield the lowest Gibbs energy for the system, along with equilibrium phase compositions and states of order.

Three main Gibbs energy minimization approaches are commonly used in the geosciences. MELTS and pMELTS (Asimow & Ghiorso, 1998; Ghiorso, 1983, 1985; Ghiorso & Sack, 1995) use Taylor series expansion to express the Gibbs energy of the system and minimize the resulting system of constrained linear equations using least squares methods. However, the thermodynamic data sets hard-wired into MELTS and pMELTS are relatively limited in their application, as they are not appropriate for handling subsolidus equilibria nor equilibria involving amphibole or biotite. Theriak-Domino uses a combination of linear programming and nonlinear local optimization methods (de Capitani & Brown, 1987) to compute the phase equilibria. Perple_X (Connolly, 2005) linearizes the problem by discretizing the equations of state in composition-order space and solves it using the simplex algorithm. A detailed review of Perple_X and Theriak-Domino methods is presented in Connolly (2017). While these two approaches have proven to be quite reliable and efficient in systems involving a limited number of components, their performance and reliability tends to decrease for higher dimensional systems. For Theriak-Domino the main limitation can be attributed to the absence of constraints during the rotation of the Gibbs hyperplane between the linear programming and nonlinear optimization stages. For Perple_X, discretization becomes increasingly expensive as the number of compositional components in the equations of state becomes larger. GeoPS (Xiang & Connolly, 2021) has recently been successful in combining these two approaches to provide the community with an efficient petrological program to easily compute phase diagrams. However, none of the above tools are MPI-parallelized for single point calculations, they are not designed to fully exploit high performance facilities, which constitutes a critical limitation for direct coupling with geodynamic modeling. The recent breakthroughs in modeling coupled mechanical and fluid/magma flow systems (e.g., Katz et al., 2022; Keller & Katz, 2016; Keller & Suckale, 2019; Keller et al., 2013, 2017; Rummel et al., 2020; Taylor-West & Katz, 2015; Turner et al., 2017) and the ongoing open-source movement in the community to simplify and unify modeling tools (e.g., Bezanson et al., 2017, Julia), however, highlight the need for an efficient, open-source, and fully parallel mineral assemblage modeling routine.

Here, we describe a new approach, MAGEMin (Mineral Assemblage Gibbs Energy Minimization) (https://github.com/ComputationalThermodynamics/magemin.git), which was developed to provide a minimization routine that is easily callable and fulfills several objectives. First, the package performs single point calculations at given pressure, temperature, and bulk-rock composition and finds the thermodynamically most stable assemblage in an automated manner with no requirement for a priori knowledge of the system and which is a requirement for integration with geodynamic software. Second, the package has been developed for stability, performance, and scalability in complex chemical systems.

Our Gibbs minimization approach combines discretization of the equations of state in composition space (Connolly, 1990) with linear programming (de Capitani & Brown, 1987) and extends the mass constrained Gibbs-hyperplane rotation (Piro, Banfield, et al., 2013; Piro, Simunovic, et al., 2013) method to account for the mixing on crystallographic sites that takes place in silicate mineral solid solutions. In this way, we overcome many of the drawbacks of the above-mentioned software packages. Moreover, since the method is developed around point-wise calculations, it is well-suited for parallelization on massively parallel machines and can be combined with an adaptive mesh refinement strategy. We demonstrate the effectiveness of our method by computing a series of phase diagrams using a large thermodynamic data set native to the THERMOCALC software and comparing the automatically calculated MAGEMin results with those obtained using THERMOCALC. The definition of the general terminology used in this contribution is given in Table 1 and the definition of the symbols is provided in Table 2.

Table 1
General Terminology

Term	Definition
System component	Chemically independent constituent (see Pauken, 2011). The collection of components define the number of chemical dimensions of the system. Here, we use oxides as system components spanning up to 11 dimensions: Na_2O–CaO–K_2O–FeO–MgO–Al_2O_3–SiO_2–TiO_2–Fe_2O_3–Cr_2O_3–H_2O
Pure phase	Pure phase (or stoichiometric phase), is a phase that has a fixed composition (or does not vary measurably from its ideal composition), for example, quartz (SiO_2)
End-member	End-member (or species (Kruskopf & Visuri, 2017) or phase component (Berman, 1991) or vertex of a polytope (Myhill & Connolly, 2021)) is an independent instance of a solution phase (with defined cation occupancy/vacancy on each site) that can be linearly combined with other end-members to span the complete site-occupancy space of a solution
Solution phase	A Solution phase is a mixture of end-members spanning a range of compositions for a single crystal structure (solid solution phase), a fluid or a melt. For instance, in a chemical subsystem restricted to FeO–MgO–SiO_2, the compositional space of olivine is covered by the linear mixture of fayalite (Fe_2SiO_4) and forsterite (Mg_2SiO_4) end-members. The composition of the end-members are expressed in oxide form (fayalite = 2FeO + SiO_2 and forsterite = 2MgO + SiO_2) but the substitution of Fe and Mg cations occurs in elemental form on the olivine crystallographic site M1 ($[Fe,Mg]^{M1}$ SiO_2)
Pseudosection	Pseudosection (or isochemical equilibrium phase diagram (de Capitani & Brown, 1987)) is a class of phase diagram in pressure-temperature space showing the fields of most stable phase equilibrium for a single bulk-rock composition
Solution phase model	Solution phase model (or equation of state (Powell, 1978)) aims to reproduce the energetic behavior of naturally occurring mineral, melt, and fluid phases. Depending on the complexity of the phase of interest, the related solution phase model is usually formulated using an ideal and a nonideal mixing term
Ideal mixing term	The ideal mixing term include both the mechanical mixture contribution, which is the linear combination of the standard Gibbs energy of the end-members and the configurational energy term that describes the change of energy when the mixture reacts to form a single phase (see Ganguly (2001) and Lanari and Duesterhoeft (2018) for more details)
Nonideal mixing term	Nonideal mixing term (or excess term) expresses the nonideal interaction between end-members (see Ganguly, 2001)

2. Methodology

2.1. Gibbs Energy Formulation

At fixed pressure P and temperature T, the integral Gibbs energy (J) of a multicomponent multiphase system G_{sys} (e.g., Gibbs, 1878; Spear, 1993) can be expressed by:

$$G_{sys} = \sum_{\lambda=1}^{\Lambda} \alpha_\lambda \sum_{i=1}^{N_\lambda} \mu_{i(\lambda)} p_{i(\lambda)} + \sum_{\omega=1}^{\Omega} \alpha_\omega \mu_\omega, \quad (1)$$

where Λ indicates the number of solution phases (mineral phases of variable composition), N_λ the number of end-members of solution phase λ, Ω the number of pure phases (mineral phases of fixed composition, also described as pure phases), and α_λ and α_ω are the mole fractions of solution phase λ and pure phase ω, respectively. $p_{i(\lambda)}$ is the fraction of end-member i dissolved in solution phase λ and $\mu_{i(\lambda)}$, μ_ω are the chemical potential of end-member i in solution phase λ and pure phase ω, respectively. An end-member is defined as an independent instance of a solution phase. In a given chemical system, the linear combination of the end-members span the complete crystallographic site-occupancy space of the solution phase.

The chemical potential of a phase is either a constant for a stoichiometric phase (Spear, 1993)

$$\mu_\omega = g^0_\omega, \quad (2)$$

or a function for dissolved end-members within a solution phase (see Ganguly, 2001, for a review)

Table 2
Symbols Definition

Symbol	Unit	Definition
R	J mol^{-1} K^{-1}	Ideal gas constant
T	K	Temperature
C	–	Total number of chemical components (oxides) in the system
F	–	Number of degrees of freedom (Gibbs-Duhem rule)
J	–	Oxide
b_j	–	Bulk-rock composition of oxide j
Φ	–	Total number of active phases
Λ	–	Total number of active solution phases
Λ	–	Solution phase
N_λ	–	Number of end-members of solution phase λ
$i(\lambda)$	–	End-member i of solution phase λ
$p_{i(\lambda)}$	mol%	Fraction of end-member i in phase λ
$x_{i(\lambda)}$	–	Penalty formulation for PGE stage of end-member i in phase λ
$x_{k(\lambda)}$	–	Compositional variable k of solution phase λ
Ω	–	Total number of active pure phases
Ω	–	Pure phase
α_λ	mol%	Fraction of solution phase λ
α_ω	mol%	Fraction of pure phase ω
a_{ij}	mol	Molar composition of oxide j in end-member i
$a_{\lambda j}$	mol	Molar composition of oxide j in solution phase λ
$a_{\omega j}$	mol	Molar composition of oxide j in pure phase ω
F	–	Normalization factor
a_j	–	Number of atom per oxide j
$e_{i(\lambda)}$	–	Molar composition of end-member i in solution phase λ
ν_s	–	Number of atoms contained in mixing site s of λ
c_i	–	Normalization constant
G_λ	J	Gibbs energy of the solution phase λ
G^{lvl}	J	Gibbs energy of system during the level stage
G_{sys}	J	Total Gibbs energy of the system
Γ_j	J	Chemical potential of pure oxide j, defining Gibbs hyperplane
Γ^{lvl}	J	Set of oxide chemical potentials obtained during leveling stage
$g^0_{i(\lambda)}$	J	Gibbs energy of reference of end-member i in phase λ
$a^{id}_{i(\lambda)}$	J	Ideal mixing term
$X^s_{e_{s,i}}$	–	Site fraction of the element $e_{s,i}$ on site s in end-member i of phase λ
$g^{ex}_{i(\lambda)}$	J	Excess energy term of end-member i in phase λ
$\mu_{i(\lambda)}$	J	Chemical potential of end-member i in phase λ
$\Delta\mu_{i(\lambda)}$	J	Gibbs energy distance of end-member i in phase λ from Gibbs hyperplane
μ_ω	J	Gibbs energy of pure phase ω
P	kg m^{-3}	Density
K_b	Pa	Adiabatic bulk modulus
K_s	Pa	Elastic shear modulus
v_p	km s^{-1}	Compressional P-wave velocity
v_s	km s^{-1}	Shear S-wave velocity

$$\mu_{i(\lambda)} = g^{0}_{i(\lambda)} + RT \log\left(a^{id}_{i(\lambda)}\right) + g^{\mathrm{ex}}_{i(\lambda)}, \tag{3}$$

where R (Jmol^{-1}K^{-1}) is the ideal gas constant, T [K] is the absolute temperature, $a^{id}_{i(\lambda)}$ is the ideal mixing term, $g^{0}_{i(\lambda)}$ the Gibbs energy of reference of the pure end-member (Helgeson, 1978; Holland & Powell, 1998), and $g^{\mathrm{ex}}_{i(\lambda)}$ is the excess energy term (Holland & Powell, 2003; Powell & Holland, 1993). The ideal mixing term $a^{id}_{i(\lambda)}$ is generally defined as $a^{id}_{i(\lambda)} = p_{i(\lambda)}$ for molecular mixing or else for mixing on crystallographic sites as

$$a^{id}_{i(\lambda)} = c_i \prod_{s} \left(X^{s}_{e_{s,i}}\right)^{\nu_s} \tag{4}$$

where $X^{s}_{e_{s,i}}$ is the site fraction of the element $e_{s,i}$ that appears on site s in end-member i of phase λ, ν_s is the number of atoms contained in mixing site s of λ, and c_i is a normalization constant that ensures that $a^{id}_{i(\lambda)}$ is unity for the pure end-member i. The total Gibbs energy of solution phase λ is given by

$$G_{\lambda} = \sum_{i=1}^{N_{\lambda}} \mu_{i(\lambda)} p_{i(\lambda)}. \tag{5}$$

At equilibrium, all pure phases and dissolved end-members in a solution phase have to satisfy the Gibbs-Duhem rule (e.g., Spear, 1988, 1993)

$$\mu_{i(\lambda),\omega} = \sum_{j=1}^{C} a_{i,\omega j} \Gamma_j, \tag{6}$$

where Γ_j is the chemical potential of the pure component j. The Gibbs-Duhem rule implies that at equilibrium, the chemical potential of all end-members of a solution phase must lie on the Gibbs hyperplane defined by Γ_j. At specified pressure and temperature, the system must satisfy the Gibbs phase rule (e.g., Spear, 1988, 1993)

$$F = C - \Phi \geq 0, \tag{7}$$

where F is the number of degrees of freedom, C is the number of components (or oxides), and Φ is the total number of phases. Finally, the system must satisfy the mass balance constraint, which implies that the ratio of chemical elements supplied by the phases at their equilibrium compositions and proportions should be equal to that in the specified bulk-rock composition b_j

$$\sum_{\lambda=1}^{\Lambda} \alpha_{\lambda} \sum_{i=1}^{N_{\lambda}} a_{ij} p_{i(\lambda)} + \sum_{\omega=1}^{\Omega} \alpha_{\omega} a_{\omega j} - b_j = 0, \tag{8}$$

where a_{ij} and $a_{\omega j}$ are composition vectors for the end-member and system components j and $\alpha_{\lambda,\omega} \geq 0$.

2.2. Gibbs Energy Minimization Strategy

For any system of fixed bulk composition, pressure, and temperature conditions, the general equilibrium conditions are given by minimizing Equation 1 while satisfying the Gibbs-Duhem rule (Equation 6) and mass constraint (Equation 7). This system of equations yields an equality-constrained optimization problem that remains notoriously difficult to solve as it involves a weighted sum of objective functions unevenly spanning the dimensional space. To compute the thermodynamic equilibrium conditions we employ a two-stage algorithm. First, we obtain an initial guess using discretized solution phases (pseudocompounds) and linear programming methods (leveling, Figures 1a and 1b) and after which a local minimization of solution phases is coupled with the PGE method (Piro, Banfield, et al., 2013; Piro, Simunovic, et al., 2013) among predicted stable phases (Figure 1c).

2.2.1. Leveling Stage

The concept of leveling is to temporarily neglect the thermodynamic contribution from mixing in the solution phases (de Capitani & Brown, 1987; Kruskopf & Visuri, 2017; Piro, Banfield, et al., 2013; Piro, Simunovic, et al., 2013). As a consequence, all end-members of solution phases and stoichiometric phases in the system are initially treated as separate pure phases. This allows us to estimate the chemical potential of system components (oxides) and the proportions of the stable pure phases using linear programming methods.

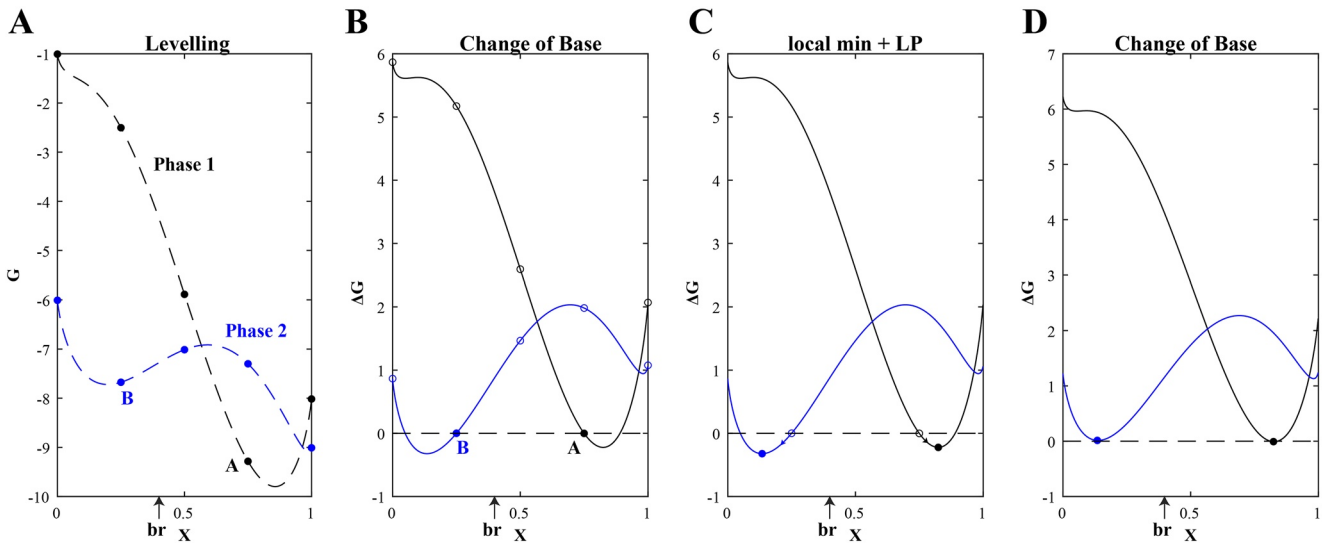

Figure 1. Simplified illustration of the minimization strategy, at pressure $P = 0$ GPa and temperature $T = 1$ K, with two solution phases λ_1 and λ_2 (modified after de Capitani and Brown (1987)). Both solution phases include two end-members of identical composition $C = [1\ 0; 0\ 1]$ (at coordinate $X = 0.0$ and 1.0). The Gibbs energy of reference of the end-members are: $G0_{i1}^{\lambda 1} = -1.0$, $G0_{i2}^{\lambda 1} = -8.0$, $G0_{i1}^{\lambda 2} = -6.0$, and $G0_{i2}^{\lambda 2} = -9.0$. The ideal mixing terms are formulated as $RT\log(x)$ with $R = 8.134$ J/mol/K and $T = 1.0$ K. The excess terms are calculated as $G_{\text{ex}}^{\lambda 1} = 35.0x_1^2x_2$ and $G_{\text{ex}}^{\lambda 2} = 35.0x_1x_2^2 + 15.0x_1^2x_2$. The total Gibbs energy of each phase is expressed as $G^{\lambda} = \sum_{i=1}^{N_{\lambda}} x_i\mu_{i(\lambda)} = \sum_{i=1}^{N_{\lambda}} x_i\,(G0_i + RT\log(x_i) + G_{\text{ex}})$. (a) During the leveling stage the G^{λ} function of each phase is discretized in composition space with a step of 0.25. Then, linear programming is used to find the combination of discrete points by minimizing the Gibbs energy of the system $G_{\text{sys}} = G^{\lambda 1} + G^{\lambda 2}$ while satisfying the mass constraint $br = [0.6\ 0.4]$, resulting in points A ($\alpha = 0.7$) and B ($\alpha = 0.3$). The Gibbs hyperplane passing through discrete phases A and B is computed by solving $\Gamma_j = \mathbf{A}^{-1}b$, where $\mathbf{A}$ is the stoichiometry matrix ([0.25 0.75; 0.75 0.25]) and b is the vector of Gibbs energy of discrete points A and B ([−9.2846; −7.6753]). (b) The whole system is rotated by recalculating the Gibbs energy curves G^{λ} as $\Delta G^{\lambda} = \sum_{i=1}^{N_{\lambda}} x_i \left(G0_i - \sum_{j=1}^{C} \Gamma_j a_{ij} + RT\log(x_i) + G_{\text{ex}}\right)$, where $\Gamma_j = [-6.8706; -10.0893]$ and a_i is the compositional vector of end-member i. This step effectively levels the plane passing through points A and B to the horizontal, that is, $\Delta G^{\lambda A,B} = 0.0$. (c) Starting from discrete points A and B (empty black and blue circles) a gradient-based method is used to find the minimum of phases λ_1 and λ_2 ($X_1 = 0.8242$ and $X_2 = 0.1345$). (d) Using the minimized points the Gibbs plane is rotated again ($\Delta\Gamma_j = [-0.3631; -0.1806]$, $\Gamma_j = \Gamma_j + \Delta\Gamma_j$) and for this simplified case study, the system is considered to have converged, as there are no phases left for which $\Delta G < 0$. The phase fractions are retrieved as $\alpha_{\lambda 1,2} = \mathbf{A}^{-1}br$ ($\alpha_{\lambda 1} = 0.3850$ and $\alpha_{\lambda 2} = 0.6151$). Note that in our formulation, unlike in de Capitani and Brown (1987), the update of the Gibbs hyperplane defined by $\Delta\Gamma_j$ is achieved using the PGE approach (Equations 26–28) modified for mixing on crystallographic sites.

Given this set of artificial "pure phases," the first step of the leveling stage minimizes

$$G^{\text{lvl}} = \min\left(\sum_{\phi=1}^{\Phi} \alpha_{\phi} g_{\phi}^{0}\right), \tag{9}$$

where Φ is the number of active phases equal to the number of system components C, α_{ϕ} is the fraction of phase ϕ, and g_{ϕ}^{0} is the chemical potential of phase ϕ and subject to the mass balance constraint

$$\sum_{\phi=1}^{\Phi} \alpha_{\phi} a_{\phi j} - b_j = 0, \tag{10}$$

where $a_{\phi j}$ is the composition vector of the phase ϕ and $\alpha_{\phi} \geq 0$. Equations 9 and 10 are solved using the linear programming method adapted from de Capitani and Brown (1987) with a special case of the simplex method (Dantzig, 1963). Upon convergence, the chemical potential of the system components defining the Gibbs hyperplane are retrieved such as

$$\Gamma^{\text{lvl}} = A^{-1}G^{\text{lvl}}, \tag{11}$$

where A is the stoichiometry matrix of the predicted stable pure phases and G^{lvl} is the Gibbs energy vector of the same set of pure phases.

During the second step of the leveling stage, solution phases are discretized (pseudocompound) and only the pseudocompounds located close or below the Gibbs hyperplane defined by Γ^{lvl} are further considered for a

second round of linear programming. The distance of a pseudocompound with respect to the Gibbs hyperplane is calculated as

$$\Delta G_\lambda = \sum_{j=1}^{C} a_{\lambda j}\Gamma_j^{\text{lvl}} - G_\lambda, \tag{12}$$

where $a_{\lambda j}$ is the composition and G_λ is the Gibbs energy of the pseudocompound.

Likewise, the distance from the Gibbs hyperplane can be calculated for each end-member dissolved in a solution phase from

$$\Delta\mu_{i(\lambda)} = \mu_{i(\lambda)} - \sum_{j=1}^{C} a_{\lambda j}\Gamma_j = g^0_{i(\lambda)} + RT\log\left(a^{id}_{i(\lambda)}\right) + g^{\text{ex}}_{i(\lambda)} - \sum_{j=1}^{C} a_{\lambda j}\Gamma_j. \tag{13}$$

Cycling through the list of pseudocompounds is achieved until no remaining pseudocompound is left with a negative $\Delta G_\lambda \leq -10^{-6}$. The leveling stage is then successfully terminated and the PGE stage is initiated.

2.2.2. Partitioning Gibbs Energy Stage

The PGE approach (e.g., Kruskopf & Visuri, 2017; Piro, Banfield, et al., 2013; Piro, Simunovic, et al., 2013) has the objective to partition the Gibbs energy of the system among the system components (i.e., Γ_j) by expressing the end-member fraction of the mass balance (Equation 8) as a function of the chemical potential of end-members (Equation 3). The key advantage of this approach is that a change in Γ_j (which we attempt to find) is directly coupled to the composition of the system, which helps the optimization process.

For a nonideal solution model where the ideal mixing term only depends on the end-member fraction (and not the site-fractions), Piro, Banfield, et al. (2013) and Piro, Simunovic, et al. (2013) first expresses the chemical potential of the end-members as a function of the Gibbs-Duhem relation by substituting Equation 6 in Equation 3 yielding

$$\sum_{j=1}^{C} a_{i(\lambda),j}\Gamma_j = g^0_{i(\lambda)} + RT\log\left(x_{i(\lambda)}\right) + g^{\text{ex}}_{i(\lambda)}. \tag{14}$$

Solving this for $x_{i(\lambda)}$ gives

$$x_{i(\lambda)} = \exp\left(\left(\sum_{j=1}^{C} a_{\lambda j}\Gamma_j - g^0_\lambda - g^{\text{ex}}_\lambda\right)/(RT)\right), \tag{15}$$

which is a direct expression of the end-member fraction $x_{i(\lambda)}$ as a function of its chemical potential and the chemical potential of the pure components of the system Γ_j. This expression is then substituted into the mass balance equation (Equation 8) yielding a set of equations (one per component) in the PGE form:

$$\sum_{\lambda=1}^{\Lambda}\alpha_\lambda\sum_{i=1}^{N_\lambda} a_{ij}\exp\left(\left(\sum_{j=1}^{C} a_{\lambda j}\Gamma_j - g^0_\lambda - g^{\text{ex}}_\lambda\right)/(RT)\right) + \sum_{\omega=1}^{\Omega}\alpha_\omega a_{\omega j} - b_j = 0, \tag{16}$$

This formulation has proven to be very successful in large chemical systems involving as many as 118 components (Piro, 2011) and relatively simple ideal and nonideal solution models (Kruskopf & Visuri, 2017; Piro, Banfield, et al., 2013; Piro, Simunovic, et al., 2013), allowing to model for the first time the temporal and spatial evolution of coupled thermochemical and nuclear reactions of irradiated fuel (Piro, Banfield, et al., 2013; Piro, Simunovic, et al., 2013).

However, this formulation cannot be directly applied to more complex solid solutions, in which mixing-on-sites must be considered, yielding an ideal entropy term that must be written in terms of site fractions (Equation 4). To extend the PGE approach to account for solution models involving site-fractions, we expand the ideal mixing term as

$$RT\log\left(a^{id}_{i(\lambda)}\right) = RT\log\left(x_{i(\lambda)}\right) + RT\log\left(\frac{a^{id}_{i(\lambda)}}{p_{i(\lambda)}}\right), \tag{17}$$

where $x_{i(\lambda)} = p_{i(\lambda)}$, which gives using Equation 3 and the Gibbs-Duhem relation of Equation 6

$$RT \log\left(x_{i(\lambda)}\right) = \sum_{j=1}^{C} a_{i(\lambda)j}\Gamma_j - g^0_{i(\lambda)} - RT \log\left(\frac{a^{id}_{i(\lambda)}}{p_{i(\lambda)}}\right) - g^{\text{ex}}_{i(\lambda)}. \tag{18}$$

Developing the log term of the right side of Equation 18 gives

$$RT \log\left(x_{i(\lambda)}\right) = \sum_{j=1}^{C} a_{i(\lambda)j}\Gamma_j - g^0_{i(\lambda)} - RT \log\left(a^{id}_{i(\lambda)}\right) - g^{\text{ex}}_{i(\lambda)} + RT \log\left(p_{i(\lambda)}\right), \tag{19}$$

which can be simplified using Equation 13 as

$$\log\left(x_{i(\lambda)}\right) = -\frac{\Delta\mu_{i(\lambda)}}{RT} + \log\left(p_{i(\lambda)}\right), \tag{20}$$

and rearranged as

$$x_{i(\lambda)} = \exp\left(-\frac{\Delta\mu_{i(\lambda)}}{RT}\right) p_{i(\lambda)}, \tag{21}$$

where $x_{i(\lambda)}$ is the expression for end-member fraction used in the subsequent PGE formulation and $p_{i(\lambda)}$ is the end-member fraction as computed by the solution phase model. After the leveling stage, and as long as the Gibbs-Duhem constraint is not respected, $\Delta\mu_{i(\lambda)} \neq 0$ and $x_{i(\lambda)} \neq p_{i(\lambda)}$. During the course of the PGE iterations, the term $\exp\left(-\frac{\Delta\mu_{i(\lambda)}}{RT}\right)$ tends to 1.0 as $\Delta\mu_{i(\lambda)}$ tends to 0.0, which enforces that at convergence the chemical potential of all the endmembers of predicted stable solution phases lie on the Gibbs hyperplane (Gibbs-Duhem rule) and that $x_{i(\lambda)} = p_{i(\lambda)}$. The exponential dependency of $x_{i(\lambda)}$ on $\mu_{i(\lambda)}$ imposes the Gibbs hyperplane computed during leveling to be sufficiently close to solution to ensure convergence.

Equation 21 is then substituted in Equation 8 yielding

$$\sum_{\lambda=1}^{\Lambda} \alpha_\lambda \sum_{i=1}^{N_\lambda} a_{ij} x_{i(\lambda)} + \sum_{\omega=1}^{\Omega} \alpha_\omega a_{\omega j} - b_j = 0, \tag{22}$$

which has the advantage that it effectively couples the mass balance constraint (Equation 8) and the chemical potential of pure components (Equation 6) (Kruskopf & Visuri, 2017; Piro, Banfield, et al., 2013; Piro, Simunovic, et al., 2013). Additionally, the sum of the end-member fractions of a solution phase must equal unity at convergence, that is,

$$\sum_{i=1}^{N_\lambda} x_{i(\lambda)} - 1 = 0, \tag{23}$$

and the stoichiometric phases must lie on the Gibbs hyperplane, that is,

$$\sum_{j=1}^{C} a_{\omega j}\Gamma_j - g^0_\omega = 0. \tag{24}$$

This results in the following system of equations

$$f_v = \sum_{\lambda=1}^{\Lambda} \alpha_\lambda \sum_{i=1}^{N_\lambda} a_{ij} x_{i(\lambda)} + \sum_{\omega=1}^{\Omega} \alpha_\omega a_{\omega j} - b_j, \tag{25}$$

$$h_l = \sum_{i=1}^{N_\lambda} x_{i(\lambda)} - 1, \tag{26}$$

$$q_k = \sum_{j=1}^{C} a_{\omega j}\Gamma_j - g^0_\omega. \tag{27}$$

Equations 25–27 are solved using a Newton-Raphson approach such as

$$\mathbf{J}\Delta y = -F, \tag{28}$$

where $\mathbf{J}$ is the Jacobian of the system of equations f_v, h_l, and q_k expressed as

$$\mathbf{J} = \begin{pmatrix} \frac{\partial f_v}{\partial \Gamma_j} & \frac{\partial f_v}{\partial \alpha_\lambda} & \frac{\partial f_v}{\partial \alpha_\omega} \\ \frac{\partial h_l}{\partial \Gamma_j} & \frac{\partial h_l}{\partial \alpha_\lambda} & \frac{\partial h_l}{\partial \alpha_\omega} \\ \frac{\partial q_k}{\partial \Gamma_j} & \frac{\partial q_k}{\partial \alpha_\lambda} & \frac{\partial q_k}{\partial \alpha_\omega} \end{pmatrix} = \begin{pmatrix} \sum_{\lambda=1}^{\Lambda} \alpha_\lambda \sum_{i=1}^{N_\lambda} x_{i(\lambda)} a_{ij} a_{iv} & \sum_{i=1}^{N_\lambda} x_{i(\lambda)} a_{ij} & a_{\omega j} \\ \sum_{i=1}^{N_\lambda} x_{i(\lambda)} a_{ij} & 0 & 0 \\ a_{\omega j} & 0 & 0 \end{pmatrix}, \tag{29}$$

F is the residual vector defined as

$$F = (f_1, \ldots, f_C, h_1, \ldots, h_\Lambda, q_1, \ldots, q_\Omega), \tag{30}$$

and Δy is the set of variable we solve for

$$\Delta y = (\Delta\Gamma_1, \ldots, \Delta\Gamma_C, \Delta\alpha_1, \ldots, \Delta\alpha_\Lambda, \Delta\alpha_1, \ldots, \Delta\alpha_\Omega). \tag{31}$$

At the beginning of a PGE iteration, all considered solution models are minimized while taking inequality constraints into account. In MAGEMin, we employ the optimization library NLopt (Johnson, 2021) and the gradient-based CCSAQ algorithm (Svanberg, 2002). This algorithm supports inequality constraints which is a requirement to minimize the solution models as the ideal mixing term is a function of the site-fractions, which have to be ≥0. An example of a solution model derivation for a gradient-based inequality constraint optimization is given in the Appendices.

Subsequently, Equation 28 is solved and the set of variables is updated as $y^1 = y^0 + \tau\Delta y$ where τ is an under-relaxing factor defined as

$$\tau = \min\left(\frac{\frac{0.025}{\delta}}{\left\| \alpha_{\lambda,\omega}^{\max} \right\|_2^2}, \frac{\frac{2.5}{\delta}}{\left\| \Gamma_j^{\max} \right\|_2^2}, 1.0 \right), \tag{32}$$

where

$$\delta = 192.0 e^{\left[-8.0 \cdot \left(\left\| \Delta b_j \right\|_2^2 \right)^{0.26} \right]} - 1.0, \tag{33}$$

is an inner under-relaxing factor linked to the residual norm of the mass constraint. Effectively, δ decreases the maximum allowed step size of $\Delta\Gamma_j$ and $\Delta\alpha_{\lambda,\omega}$ when the norm of the mass constraint decreases. δ has been optimized in a manual manner across the pseudosections presented in this study to increase the performances and stability of the computation. We choose this option because the use of backtracking line search methods has proven to be rather inefficient as, to converge, the system has to be able to temporarily relax constraints. Although the current definition of δ makes it a proud member of the family of "magic" numbers, we find that the minimization results remain quite consistent as long as the relaxing factor is small enough.

During the iterations, a phase is removed from the active assemblage when its fraction is ≤0.0 and a phase is added when its driving force ΔG_λ is ≤0.0, that is, the phase has a lower energy than/or that which is lying on the Gibbs hyperplane.

The system is considered to have converged when the norm of the mass balance residual vector, the residual of the sum of the end-member fractions, and the driving force of the solution phases are lower than 10^{-5}.

2.2.3. Solution Phase Solvi

Solvi are regions of unmixing within a solution. They can be detected when a stable or metastable assemblage contains two distinct phases with the same structure but different composition, such as coexisting augite and pigeonite for clinopyroxene (e.g., Gasparik, 2014). Computationally, the two phases are represented by local

minima at two different compositions within a single isobaric-isothermal G-surface of an equation of state. They are handled in several ways.

Firstly, after the leveling stage, if multiple discretized points on the G-surface of the same solution phase are predicted in the stable mineral assemblage, they are initially all treated as potential solvi candidates. Subsequently, they are merged after the local minimization step if they converge to the same local minimum $\left(\|\Delta x_k\|_2^2 \leq 10^{-2}\right)$.

Secondly, during the course of the PGE iterations a solution phase can be duplicated and checked for solvi when its active set of compositional variables is too far from its starting values, that is,

$$\left\|x_\lambda^t - x_\lambda^0\right\|_2^2 \geq x_{\text{step}}\sqrt{N_x}, \tag{34}$$

where x_λ^t is the actual set of compositional variables, x_λ^0 is the initial set of compositional variables, N_x is the number of compositional variables of solution phase λ, and x_{step} is the discretization step of the solution phase λ.

Finally, when getting close to solution $\Delta b_j \leq 10^{-4}$ discretized points of solution phases close to the Gibbs hyperplane but compositionally away from the active solution phase are locally minimized. In the event the driving force of a tested point $\Delta G_\lambda^{\text{tested}} \leq 0.0$, the point is added to the system. The latter strategy ensures that solvi are not overlooked and that the system converges toward global minimum (no phase lies below the Gibbs hyperplane).

2.2.4. Failed Minimization Contingency Plan

While local minimization using NLopt (Johnson, 2021) and CCSAQ algorithm (Svanberg, 2002) has proven to be quite efficient and reliable, in some cases the site-fraction inequality constraints can be slightly violated, which lead to wrong values of ΔG_λ and $\Delta\mu_{i(\lambda)}$ and therefore to divergence of the overall algorithm.

To avoid this, site-fractions are tested after every local minimization of solution phases and in the event a site-fraction is violated, the set of compositional variables is brought back to the feasible domain using the null-space formulation described in Feppon et al. (2020) such as

$$\Delta x_k = -\alpha_c\left(\mathbf{G}^T\left(\mathbf{G}\mathbf{G}^T\right)^{-1} g\right), \tag{35}$$

where Δx_k is the compositional variable step toward the feasible domain, g is a vector of violated site-fraction constraints, $\mathbf{G}$ is the Jacobian of the violated site-fractions, and $\alpha_c = 0.1$ is an under-relaxing factor. This approach proved to be robust and the solution phase is generally brought back into the feasible domain within 4–5 iterations.

In the event, convergence is not achieved using the default tolerance, the tolerance can be relaxed by up to one order of magnitude ($\leq 2 \times 10^{-4}$). If convergence is still not obtained, the minimization is considered to have failed. In all cases a code, MAGEMin sends back the status of the minimization (0, success; 1, relaxed tolerance; and 2, failure).

2.3. Data Set Implementation

To improve performance and benchmark the results with THERMOCALC, the thermodynamic data set used natively in THERMOCALC was translated directly into C routines for MAGEMin and implemented without transformation of variables or coordinate systems. This eliminates inconsistencies and minimizes the risk of introducing mistakes. Appendix A gives an overview of equation of state construction in the thermodynamic data set.

2.4. Normalization for Mass Balance

Like THERMOCALC, MAGEMin accepts input bulk compositions expressed in terms of normalized numbers of oxide units (SiO_2, Al_2O_3, CaO, …). However, the phases present in the system at equilibrium will in general be written on a variety of formula units (e.g., (Mg, Fe, Ca)(Mg, Fe)SiO_4, (K, Na, Ca, Mg, Fe)(Mg, Fe, Al,Fe^{3+}, Cr)(Si, Al)$_2O_6$, …). To be able to compare the amounts of phases present in a meaningful way, MAGEMin follows THERMOCALC in expressing the amounts of phases present on a 1-atom basis. The Gibbs energies of phases must therefore be normalized.

The normalized Gibbs energy of a model solution phase is expressed as

Table 3
Bulk-Rock Compositions (mol%) Used to Produce the Igneous Phase Diagrams

Name	SiO_2	Al_2O_3	CaO	MgO	FeO	K_2O	Na_2O	TiO_2	O	Cr_2O_3	H_2O
Demo	70.69	16.63	4.56	–	–	4.45	3.67	–	–	–	–
KLB1 peridotite	38.49	1.776	2.824	50.57	5.89	0.01	0.25	0.10	0.096	0.109	–
RE46 basalt	50.72	9.16	15.21	16.25	7.06	0.01	1.47	0.39	0.35	0.01	–
Tonalite 101	66.01	11.98	7.06	4.16	5.30	1.57	4.12	0.66	0.97	0.01	50.0
Water-bearing basalt	50.08	8.69	11.67	12.14	7.78	0.22	2.49	1.00	0.47	0.01	5.44
MIX1G pyroxenite	45.25	8.89	12.22	24.68	6.45	0.03	1.39	0.67	0.11	0.012	–
N-MORB basalt	53.21	9.41	12.2 1	12.21	8.65	0.09	2.90	1.21	0.69	0.02	–

Note. Note that for readability purpose the Tonalite 101 bulk-rock composition is not displayed normalized to 100.

$$fG_\lambda = f\sum_{i=1}^{N_\lambda}\mu_{i(\lambda)}p_{i(\lambda)}, \tag{36}$$

where f is the normalization factor defined as

$$f = \frac{\sum_{j=1}^{C} b_j a_j}{\sum_{i=1}^{N_\lambda} p_{i(\lambda)} \sum_{j=1}^{C} e_{i(\lambda)j} a_j}, \tag{37}$$

where a_j is the number of atom per oxide and e_i is the molar composition of end-member i of solution phase λ.

The first derivative of the objective function (Equation 5), necessary to conduct gradient-based minimization, is computed using the chain rule as

$$\frac{\partial f G_\lambda}{\partial x_k} = \left(\mu_{i(\lambda)} - \frac{\sum_{j=1}^{C} e_{i(\lambda)j} a_j}{\sum_{l=1}^{N_\lambda}\sum_{j=1}^{C} e_{l(\lambda)j} a_j} G_\lambda\right) f \frac{\partial p_{i(\lambda)}}{\partial x_k}, \tag{38}$$

where x_k are the compositional variables of solution phase λ.

2.5. Solution Phase Discretization

The set of pseudocompounds used during the first leveling stage (Section 2.2.1) and tested when getting close to convergence is precomputed using a python Jupyter-Notebook and implemented as C functions to improve performance. The discretization step for each solution phase is chosen to be $0.05 \leq \Delta x_k^{\text{step}} \leq 0.33$ such that the total number of discrete points per solution phase ranges between 100 and 6,000 depending on the number of compositional variables (dimensionality). The currently used compositional variable steps for the discretization of the solution phases are spn, 0.199; bi, 0.124; cd, 0.098; cpx, 0.249; ep, 0.049; g, 0.198; hb, 0.329; ilm, 0.049; liq, 0.198; mu, 0.198; ol, 0.098; opx, 0.249; and pl4T, 0.049; as fl is largely dominated by water, we only use one pseudocompound made up of 100% of the pure water endmember.

3. Algorithm Demonstration

To demonstrate how the extended PGE algorithm compares to a linear programming (Theriak-Domino) approach, we present a simplified application in the Na_2O–CaO–K_2O–Al_2O_3–SiO_2 (NCKAS) chemical system. This application includes two pure phases, sillimanite and quartz, and activity-composition (a–x) relations for feldspar (pl4T) (Holland et al., 2021). The bulk-rock composition used in this example is presented in Table 3 as "demo" and the pressure and temperature conditions are fixed at 600°C and 0.3 GPa.

The results of the Gibbs energy minimization are shown in Figure 2. At equilibrium the Gibbs energy of the system is 1080.8358 J and the mineral assemblage is characterized by quartz (8.123 mol%), sillimanite (9.614 mol%), and two feldspar (41.084 and 41.179 mol%) (Figure 2). Although both the LP and the PGE methods give very

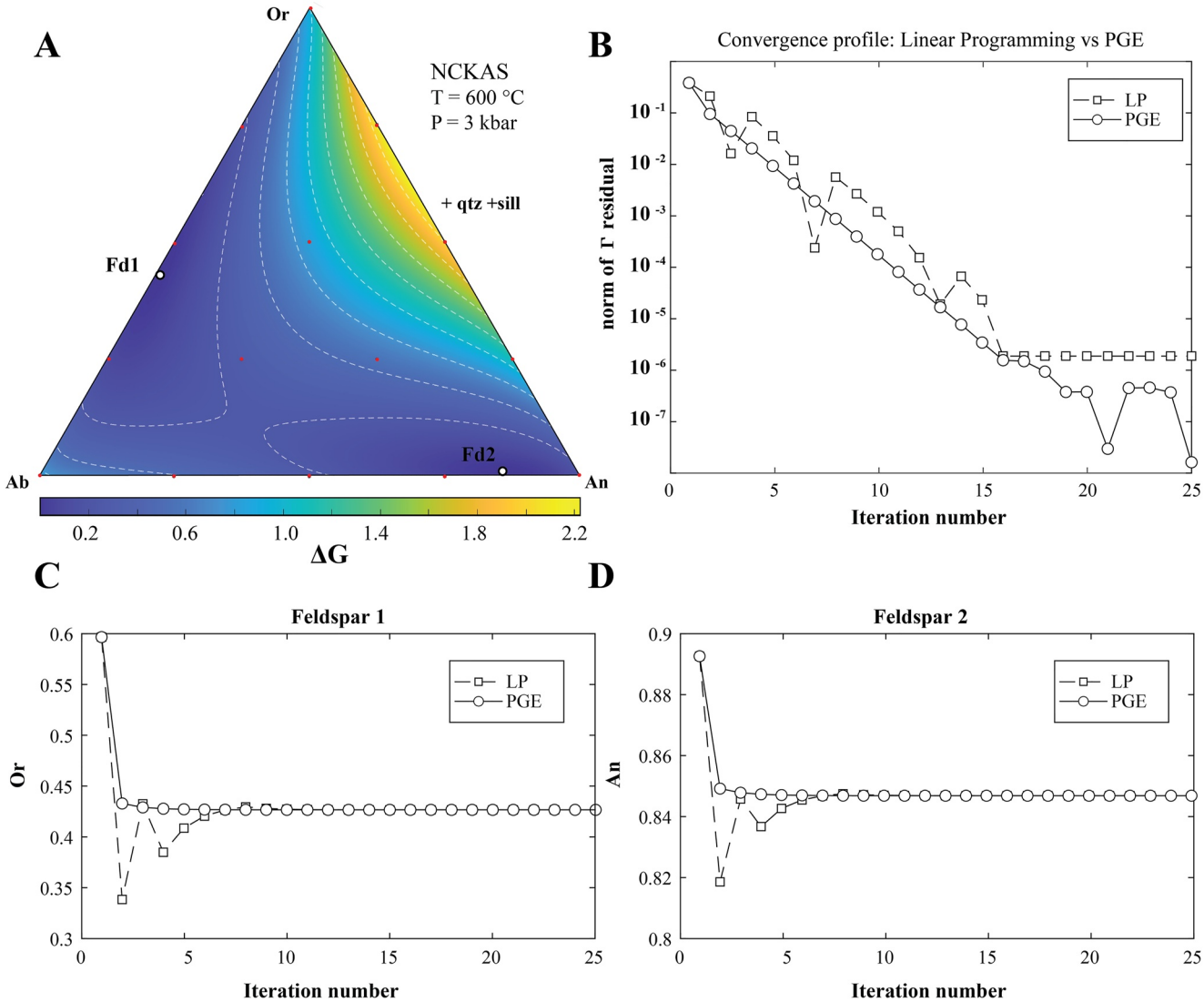

Figure 2. Comparison of Linear Programming (LP) versus the extended Partitioning Gibbs Energy (PGE) approach. (a) ΔG energy of the ternary feldspar at equilibrium. "Fd1" and "Fd2" are the two feldspar phases coexisting at equilibrium. The red dots represent the starting set of discretized points of the feldspar solution model for both LP and PGE approach. (b) Comparison of the convergence profile between LP and extended PGE approaches. (c) Orthose content of feldspar 1 as a function of the minimization iteration. (d) Anorthite content of feldspar 2 as a function of the minimization iteration.

similar results and both exhibit superlinear convergence (Figure 2b), important differences can be observed. First, the residual on the chemical potential of the system components $\|\Gamma_j\|_2^2$ is, at convergence, one order of magnitude lower with the PGE method (Figure 2b). Note that the absolute accuracy is controlled by the tolerance of the nonlinear optimizer. Here 10^{-10} was used with fmincon MATLAB solver. Secondly, on a log10 basis, the PGE convergence profile is piecewise linear while the LP profile exhibits significant oscillations (Figure 2b). The LP oscillations are caused by under/overshooting during local minimization (Figures 2c and 2d) which is not observed for the PGE approach (Figures 2c and 2d).

This key difference in convergence behavior between the LP and the PGE methods is related to how the Gibbs hyperplane is rotated after the gradient-based minimization stage. For the LP method, rotation is achieved by using the Gibbs energy of the minimized points irrespective of the chemical potential of their constitutive endmembers (see Equation 9). Instead, for the extended PGE approach, the rotation is conducted by solving the mass constraint equation where the endmember fractions are penalized using the chemical potentials of the endmembers (see Equation 8). The additional constraint drives the rotation of the Gibbs hyperplane in a direction that also decreases the Gibbs-Duhem residual of all species, that is, the Gibbs hyperplane of each individual

solution phase is rotated to be parallel to the system Gibbs hyperplane. This efficiently removes under/overshooting during the nonlinear stage (Figures 2c and 2d).

This example has been performed with MATLAB using the optimization toolbox and the script is available at https://github.com/ComputationalThermodynamics/sandbox.git.

4. Application to Igneous Systems

Below, we demonstrate our approach by presenting a variety of phase diagram calculations, which we compare with the output from THERMOCALC. Pseudosections map the most stable phase equilibrium to occur in a specified bulk-rock composition, as a function of pressure and temperature. The bulk compositions used in our calculations are shown in Table 3 and are defined in the systems Na_2O–CaO–K_2O–FeO–MgO–Al_2O_3–SiO_2–TiO_2–Fe_2O_3–Cr_2O_3 (NCKFMASTOCr; "dry") or NCKFMASTOCr + H_2O (NCKFMASHTOCr; "wet").

We used a thermodynamic data set based on that of Holland et al. (2018), including the minor published updates to the equations of state for solution phases current as of 23 Jan 2022 (see http://hpxeosandthermocalc.org). The updates include a change in the feldspar equation of state to that of Holland et al. (2021). We used version 6.34 of the internally consistent data set of end-member thermodynamic properties (Tomlinson & Holland, 2021). The thermodynamic data set as a whole includes equations of state for the pure stoichiometric phases quartz (q), cristobalite (crst), tridymite (trd), coesite (coe), stishovite (stv), kyanite (ky), sillimanite (sill), andalusite (and), rutile (ru), and sphene (sph). It also represents the solution phases spinel (spn), biotite (bi), cordierite (cd), clinopyroxene (cpx), orthopyroxene (opx), epidote (ep), garnet (g), hornblende (hb), ilmenite (ilm), silicate melt (liq), muscovite (mu), olivine (ol), ternary feldspar (pl4T), and aqueous fluid (fl). An outline of the construction of the thermodynamic data set is given in Appendix A. Full documentation and THERMOCALC input files for the thermodynamic data set can be downloaded from http://hpxeosandthermocalc.org, designated as an accompaniment to this paper.

4.1. Example Pseudosections

Pseudosections were computed using MAGEMin and processed using an MATLAB Graphical User Interface that employs adaptive mesh refinement similar to what is done in Perple_X (Connolly, 2005; Figure 2). The MATLAB application sends a list of pressure-temperature points to MAGEMin for a specified bulk-rock composition and receives back the stable phase mineral assemblage. MAGEMin is parallelized using MPI and can therefore take advantage of multicore processor architectures or be deployed on local or larger remote computing servers.

In total, 6 pseudosections are presented: KLB-1 peridotite (e.g., Takahashi, 1986), RE46 Icelandic basalt (Yang et al., 1996), water-oversaturated tonalite 101 (Piwinskii, 1968), wet basalt and two additional N-MOR basalt (Gale et al., 2013), and MIX1G pyroxenite (Hirschmann et al., 2003). Among them, KLB-1, RE46, wet basalt, and Tonalite 101 are directly benchmarked against pseudosections produced with THERMOCALC (Figures 3 and 4).

The pseudosections for KLB-1 peridotite (Figures 3a and 3b), RE46 Icelandic basalt (Figures 3c and 3d), Tonalite 101 (Figures 4a and 4b), and Wet basalt (Figures 4c and 4d) were computed with both THERMOCALC and MAGEMin. The pseudosections for KLB-1 and RE46 were computed in the KNCFMASTOCr system from 0 to 5 GPa and from 800°C to 2,000°C, and from 0 to 1.2 GPa and from 1,000°C to 1,400°C, respectively. The pseudosections for T101 and Wet Basalt were computed in the KNCFMASHTOCr system from 0 to 0.25 GPa and from 650°C to 925°C, and from 0 to 2.4 GPa and from 800°C to 1,400°C, respectively. For MAGEMin, the total number of minimized points per pseudosection varies from 80,000 to 100,000. The resulting 4 pseudosections produced with MAGEMin show nearly identical results to the one produced with THERMOCALC (Figures 3 and 4).

The pseudosections for N-MOR basalt and MIX1G pyroxenite were computed only with MAGEMin in the KNCFMASTOCr chemical system (Figures 4a and 4b) from 0 to 1.2 GPa and from 500°C to 1,400°C, and from 0 to 2.0 GPa and from 600°C to 1,600°C, respectively.

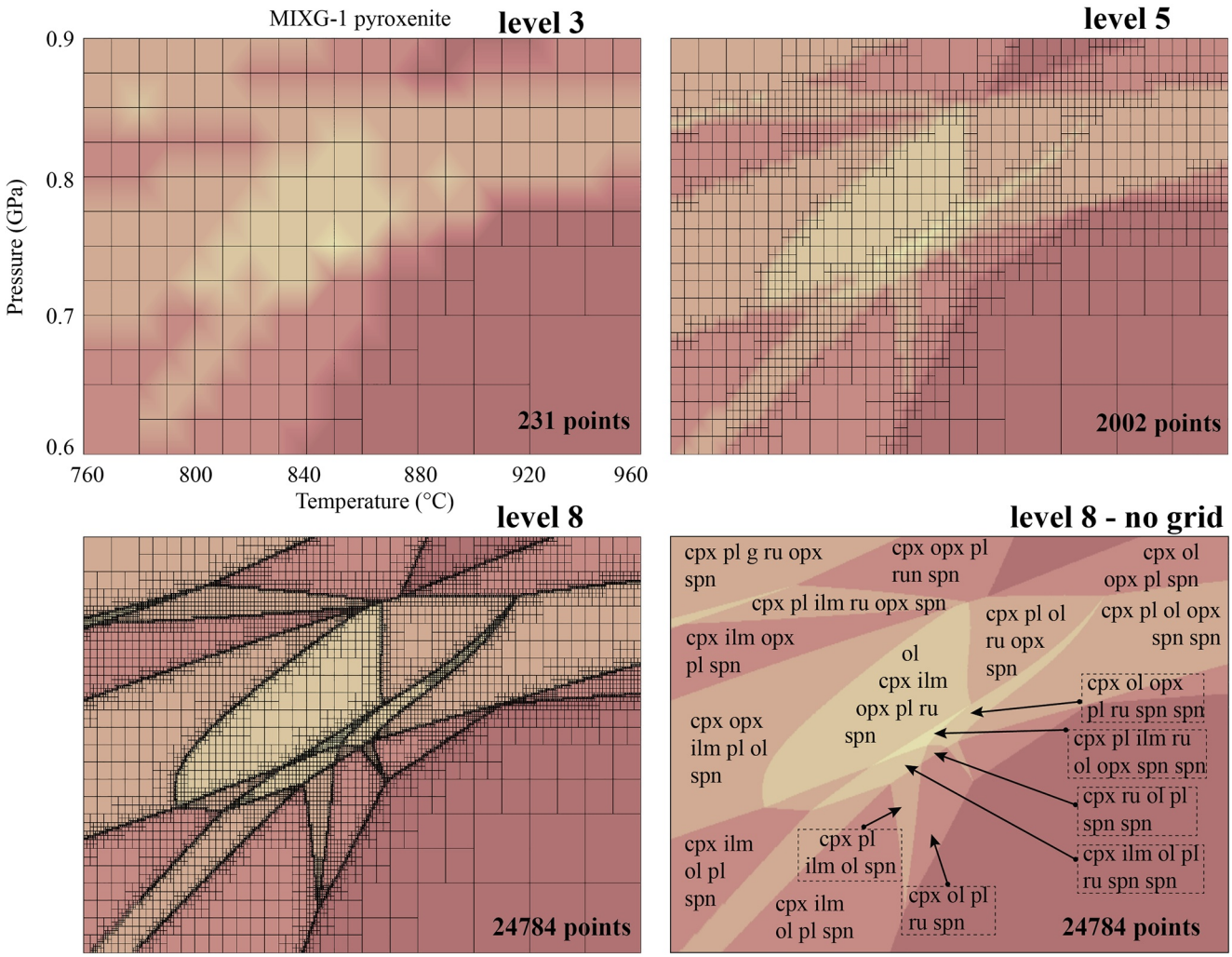

Figure 3. Illustration of the adaptive mesh refinement strategy used for pseudosection computation. Adaptive mesh refinement is illustrated for a subsection of MIX1G phase diagram displaying complex phase relations (see Figure 5b). In total eight levels of refinements are processed with an initial pressure-temperature step of 0.1 GPa and 40°C. A grid cell is refined by splitting in four smaller cells, when at least one of the four corners exhibits a different mineral assemblage. This allows us to progressively increase the resolution along reaction lines and properly resolve them as the levels of refinement increase (see level 8). Moreover, this strategy allows us to significantly reduce the total number of minimization compared to a uniform refinement reaching the same resolution, which would have required 245,760 points for this example.

4.2. Seismic Velocities

Seismic velocities (see Figure 7) are computed following the approach described in Connolly and Kerrick (2002) such as:

$$v_p = \sqrt{\frac{K_b + \frac{4K_s}{3}}{\rho}}, \tag{39}$$

and

$$v_s = \sqrt{\frac{K_s}{\rho}}, \tag{40}$$

where v_p is the P-wave velocity, v_s the S-wave velocity, ρ the density, K_b the adiabatic bulk modulus, and K_s is the elastic shear modulus. The adiabatic bulk modulus is calculated from the thermodynamic data as

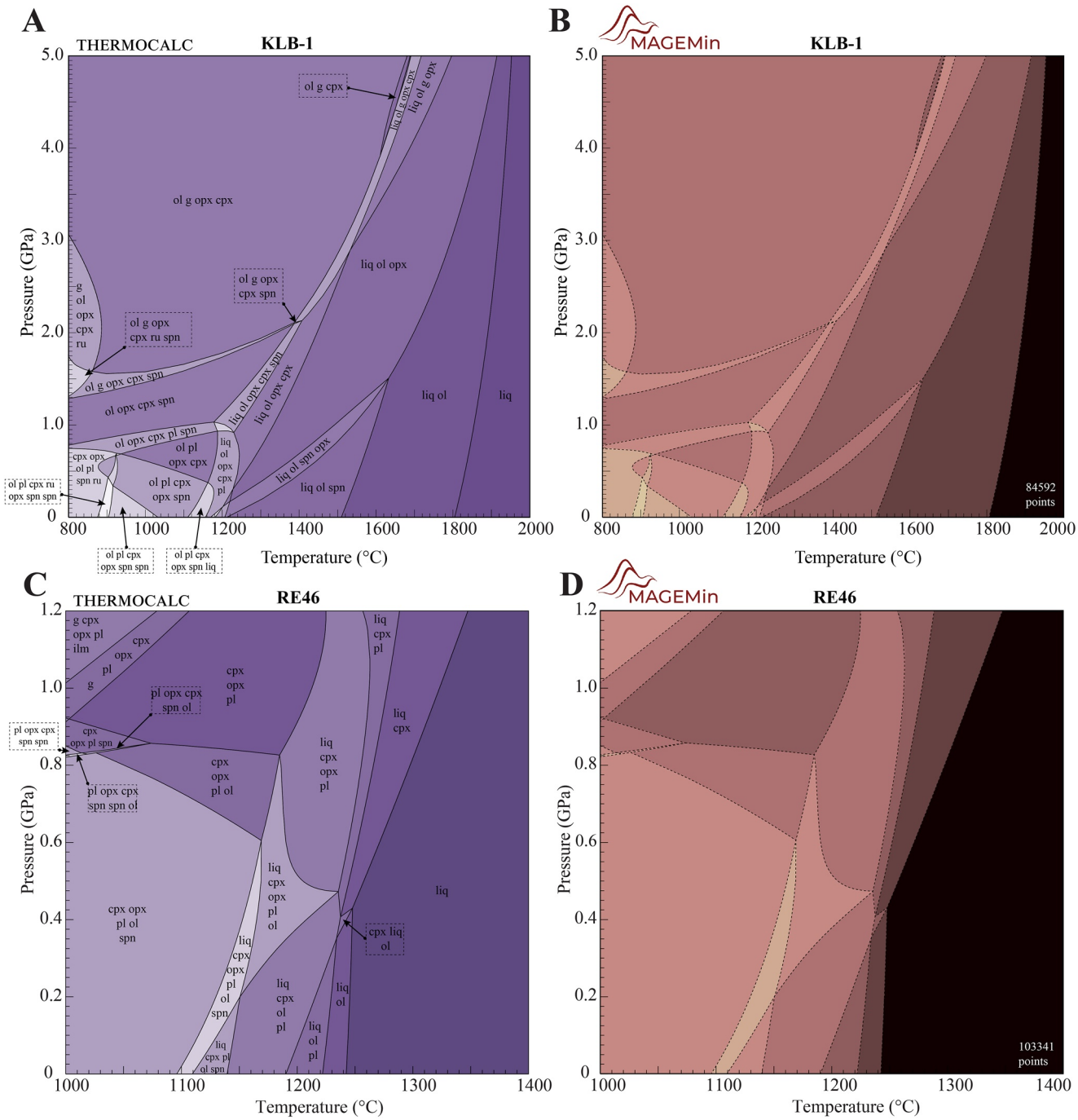

Figure 4. Comparison of dry pseudosections produced by THERMOCALC and Mineral Assemblage Gibbs Energy Minimizer (MAGEMin). (a and b) KLB-1 peridotite. (c and d) RE46 Islandic basalt. For comparison, THERMOCALC reaction lines are shown as dashed lines in the MAGEMin pseudosection. Shading represents the variance of the system.

$$K_b = -\frac{\partial G_{sys}}{\partial P}\left[\frac{\partial^2 G_{sys}}{\partial P^2} + \left(\frac{\partial}{\partial P}\frac{\partial G_{sys}}{\partial T}\right)^2 \Big/ \frac{\partial^2 G_{sys}}{\partial T^2}\right]^{-1}. \quad (41)$$

Shear moduli cannot be computed from thermodynamic data and are therefore calculated using an empirical relation (Connolly & Kerrick, 2002):

$$K_s = K_s^0 + T\frac{\partial K_s}{\partial T} + P\frac{\partial K_s}{\partial P}. \quad (42)$$

The shear moduli of the appropriate phases used in this study are taken from the database provided in Perple_X (Connolly, 2005). The database is a collection of shear moduli data mainly from Helffrich (1996) and from Karki et al. (2001), Sinogeikin et al. (2000), Bailey and Holloway (2000), and Bass (1995). The bulk-rock seismic velocities are calculated without anelasticity correction, using Voight-Reuss-Hill averaging of the velocities of the constituent phases, weighted by volume fraction (Connolly & Kerrick, 2002).

5. Discussion

5.1. Minimization Approach

Here, we present a new Gibbs energy minimization approach applied to multiphase multicomponent systems. While some of the key ideas of our approach are based on the method of PGE (e.g., Kruskopf & Visuri, 2017; Piro, Banfield, et al., 2013; Piro, Simunovic, et al., 2013) we extended it to account for modeling of mineral solid solutions involving mixing-on-sites (Equations 17–21). In Piro, Banfield, et al. (2013) and Piro, Simunovic, et al. (2013), the fraction of the end-members are updated using Equation 15, whereas in our formulation the PGE stage is used to decrease the residual between $x_{i(\lambda)}$ and $p_{i(\lambda)}$. Essentially, the expression $\exp\left(-\frac{\Delta\mu_{i(\lambda)}}{RT}\right)$ of Equation 21 forces the system to progressively satisfy the Gibbs-Duhem rule by penalizing the fraction of end-members (hence, the mass constraint) computed during local minimization. However, to avoid divergence, the PGE stage requires a good initial guess, proper set of under-relaxing factors, and more critically, a highly consistent local minimization step. Although NLopt (Johnson, 2021) implements several gradient-based minimizers with inequality constraints (SLSQP, MMA, and CCSAQ), we find that the CCSAQ algorithm (Svanberg, 2002) yields by far the best consistency and precision compared to SLSQP and linear MMA.

5.2. Consistency

The application of MAGEMin to the igneous thermodynamic data set of Holland et al. (2018) shows very good agreement with pseudosections produced with THERMOCALC (Figures 3 and 4). However, minor differences can be observed in some regions. For instance the tonalite 101 pseudosection produced with MAGEMin (Figure 4b) has an irregular reaction line at ca. 750°C and 0.1 GPa (Figure 3b). This is caused by oscillations when getting close to convergence, where a relaxed solution tolerance was accepted (tol $\leq 10^{-4}$ instead of the default value of 10^{-5}). Similar irregular reaction lines related to slightly relaxed tolerances can be observed in N-MOR basalt at ca. 750°C and 0.88 GPa (Figure 5a) and in MIX1G pyroxenite at ca. 1,250°C and 1.4 GPa (Figure 5b).

5.3. Computational Efficiency

Pseudosections presented in this work have been run in parallel on 6 logical processors on an Intel(R) Core(TM) i5-11400H. Single point calculation time has been averaged for each pseudosection that yielded 96 ms for KLB-1, 122 ms for RE46, 186 ms for Tonalite 101, and 162 ms for Wet Basalt. The large increase of calculation time for water-bearing compositions (Tonalite 101 and Wet Basalt) stems from having a larger number of discrete points during leveling (+40 ms) and a larger number of global iterations to reach convergence (+40 to 60 iterations on average).

To roughly compare our results with Perple_X (Connolly, 2005), we recomputed the KLB-1 pseudosection at similar resolution (±9,000 grid points, Figure 8). Although Perple_X did not include the last version of the data set "hp634ver.dat" at the time we generated the diagram (Figure 8a), we obtain a similar pseudosection nearly 20 times faster (Figure 8b). Note that with default option the pseudosection was computed with Perple_X in 36 min but the overall quality of the grid was quite degraded and we therefore choose to increase the grid resolution (exploratory and autorefine set to 60 and 200, respectively). In terms of single core performance, we still find that MAGEMin is nearly three times faster and yields cleaner diagrams with less visible artifacts (Figure 6). Moreover, since the current version of Perple_X is not parallelized, the computational differences are more significant in

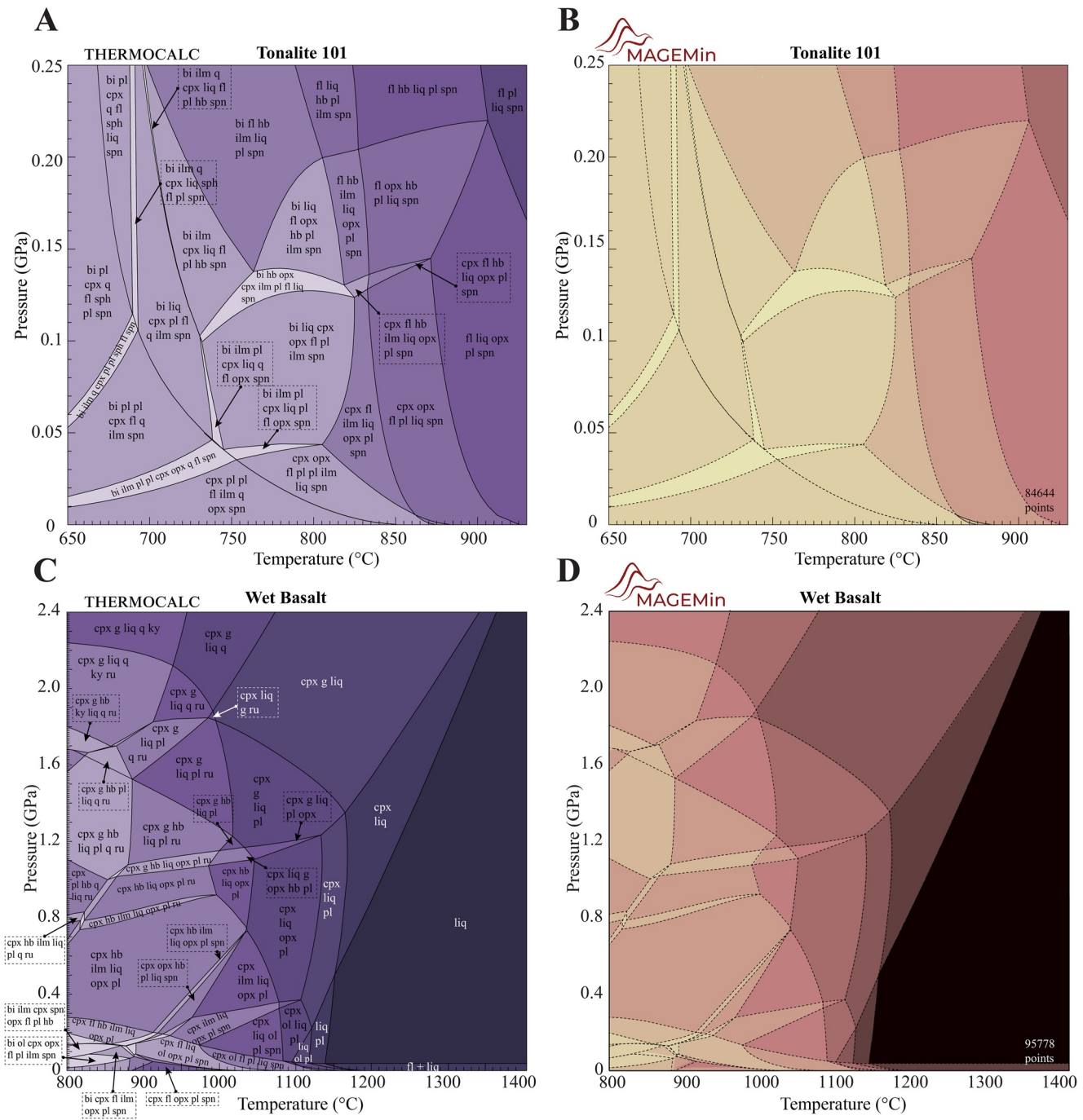

Figure 5. Comparison of wet pseudosections produced by THERMOCALC and Mineral Assemblage Gibbs Energy Minimizer (MAGEMin). (a and b) T101 tonalite. (c and d) Wet basalt.

practice. In the other comparable G-minimization software, Theriak and pMELTS, the Holland et al. (2018) thermodynamic data set is not or cannot be implemented.

5.4. Coupling With Geodynamic Codes

To facilitate coupling with geodynamic codes, we provide a Julia wrapper to MAGEMin. The Julia wrapper (called MAGEMin_C) allows the user to directly call the C functions of MAGEMin without writing data to disk first (which

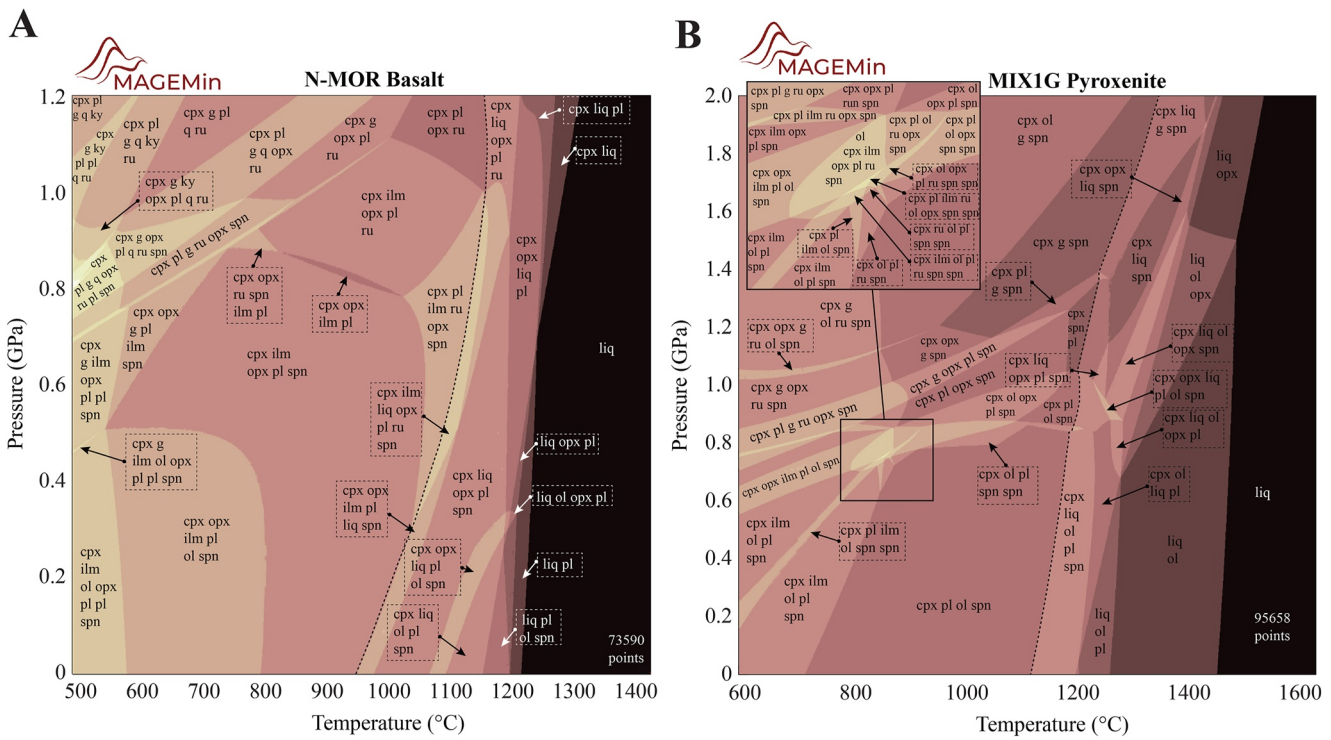

Figure 6. Pseudosections for (a) N-MOR basalt and (b) MIX1G pyroxenite.

is slow). The results of the minimization are saved into a structure that is also accessed through the Julia interface. Here, we provide a simple example.

Assuming Julia is installed, to install MAGEMin_C, first open a Julia terminal and type:

```
julia>]                                          # opens the package manager
pkg>  add  MAGEMin_C                             # MAGEMin_C
```

To compute a phase equilibrium, first leave the package manager (using backspace) and enter the following commands:

```
julia>  using  MAGEMin_C                         # load MAGEMin_C package
julia>  gv, DB       =  init_MAGEMin();          # initializes MAGEMin
julia>  P_kbar,T_C   =  8.0, 800.0;
julia>  bulk_rock    =  get_bulk_rock(gv, 0);# bulk-rock composition for test
                                                 0 (KLB-1 peridotite)
julia>  gv.verbose   =  -1;                      # switch off run-time verbose
julia>  out          =  point_wise_minimization(P_kbar,T_C, bulk_rock, gv,
                        DB);
julia>  print_info(out);                         # full display of the minimized
                                                 point.
```

A complete guide of the Julia interface is provided on the MAGEMin webpage.

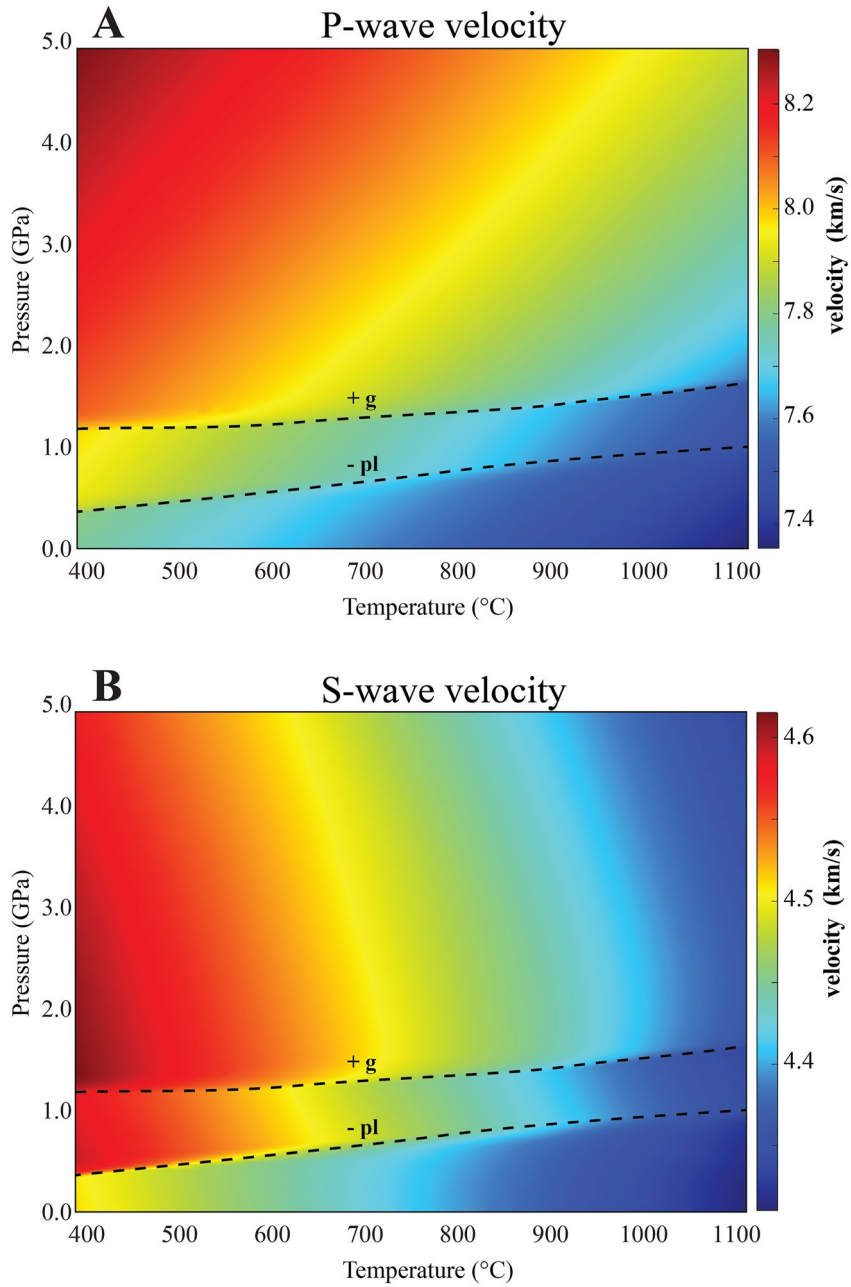

Figure 7. Example of seismic velocity computation for KLB-1 peridotite. (a) P-wave velocity. (b) S-wave velocity.

5.5. Current Limitations and Known Problems

Currently, only the thermodynamic data set for igneous systems (Holland et al., 2018) has been implemented in MAGEMin. Yet, our approach is generally applicable and should thus, in principle, work with any thermodynamic data set. To account for other petrological systems, additional data sets could be implemented, for example, relevant to metapelitic (R. W. White et al., 2014) or metabasitic systems (Green et al., 2016). We expect the performance of those cases to be at least as good as the equations of state for solution phases which are somewhat simpler.

To be successful, our implementation of the PGE method heavily relies on having good initial guesses, provided here by the leveling stage. At present, our approach tends to have difficulties to converge in some cases, mainly at subsolidus conditions and for water-bearing bulk-rock compositions (<650°C). When divergence is observed,

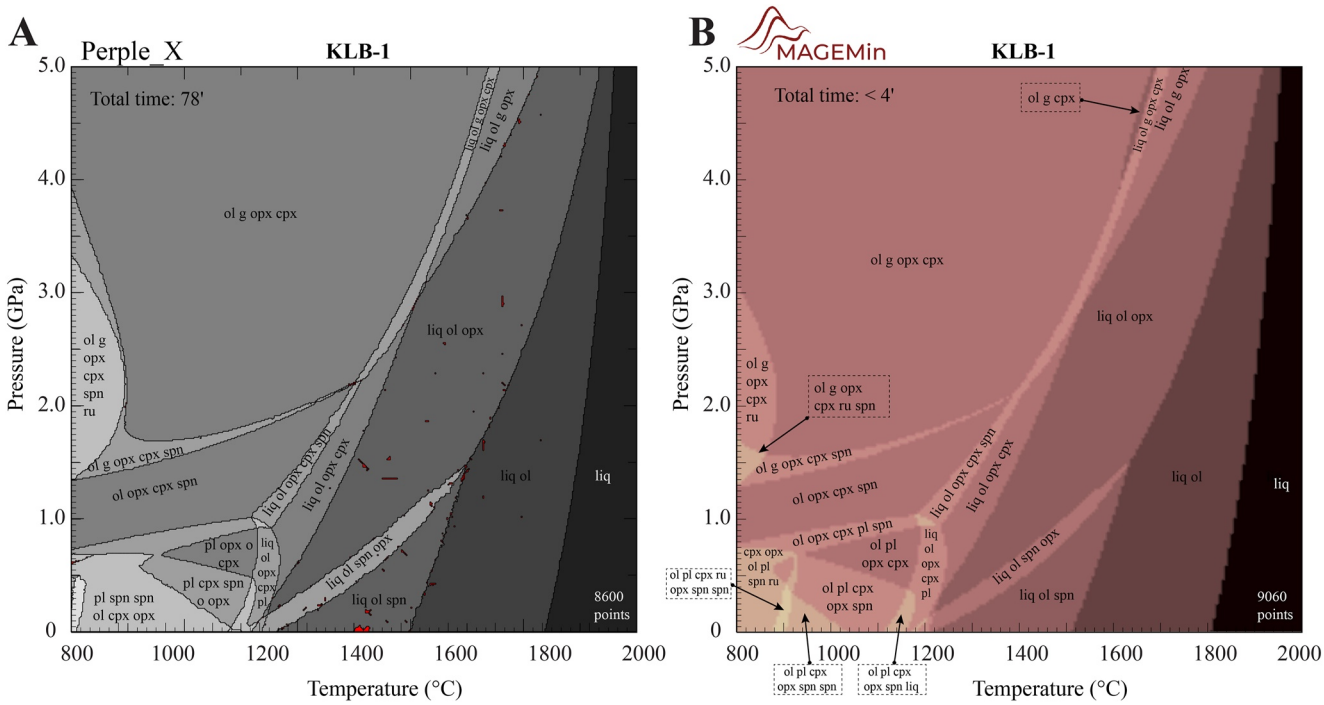

Figure 8. Comparison of KLB-1 pseudosections produced by (a) Perple_X and (b) Mineral Assemblage Gibbs Energy Minimizer (MAGEMin). For the Perple_X pseudosection, we used version 6.9.1, the database file hp633ver.dat, and the solution models Sp(HGP), Gt(HGP), Cpx(HGP), melt(HGP), O(HGP), Opx(HGP), feldspar, and excluding "enL" and "fo8L." To have a better resolution of the reaction lines, we increased the exploratory and autorefine parameters to 60 and 200, respectively. For MAGEMin, we employed four levels of grid refinement to reach a similar number of minimization points as displayed in the Perple_X log.

it usually occurs very fast and several methods are being tested to remedy that issue. One potential cause of divergence can be attributed to the current discretization approach employed during the leveling stage. Indeed, for complex solution phases, such as amphibole, the large discretization step used to keep the number of pseudocompounds relatively low (<6,000) can be a source of uneven sampling of the solution phase space, which may lead to minimization failure (likely because the minimization gets stuck in an unfeasible local minimum). A possible workaround would be to add the complete list of end-members bounding the space of valid site occupancies, following the procedure described in Myhill and Connolly (2021). A complementary option could be to precompute over a given P-T range the local minima of each solution phase and add them to the pseudocompound list.

Finally, magemin does not presently account for bulk-rock compositions that are in a different system from the set of solution phase models (equations of state) provided in Holland et al. (2018). While TiO_2, Fe_2O_3, Cr_2O_3, and H_2O can somewhat be set to 0.0, other system components cannot be ignored without reformulating some of solution phase model. However, being able to obtain stable phase equilibria in restricted chemical system is crucial to model magmatic differentiation. As a consequence, we are actively working on producing a generalized set of solution phase models accounting for reduced chemical systems.

6. Conclusions

We present a new parallel Gibbs energy minimizer that allows us to compute stable equilibria in complex multicomponent multiphase systems. We successfully extended the PGE approach to Gibbs energy functions that model mixing-on-sites and applied it to the most recent thermodynamic igneous data set. Pseudosection computation shows very consistent results with THERMOCALC and improved performance with respect to other software such as the current Perple_X version. The parallel design of MAGEMin makes it highly scalable on multicore machines. While in this contribution, we computed pseudosections using a MATLAB-based interface, MAGEMin has been developed with the objective to provide the community with a minimization package easily callable from any geodynamic codes. Such tool can also potentially provide a robust framework for thermodynamic database inversions.

Appendix A: Equations of State in the Example Thermodynamic Data Set

Here, we summarize how the equations of state are formulated in our example thermodynamic data set to that of Holland et al. (2018). We detail the information passed as input to both THERMOCALC and Mineral Assemblage Gibbs Energy Minimizer (MAGEMin), using the Holland et al. (2018) equation of state for the olivine solid solution as an example. In Appendix B, we show how MAGEMin sets up the minimization step for the model olivine solid solution.

In Holland et al. (2018) and related thermodynamic data sets such as R. W. White et al. (2014) and Green et al. (2016), the equation of state of a mineral is assembled up to four components (for a pure phase, consisting of a single end-member, only aspects 2 or 3 apply):

1. A choice about what composition space the model solution phase should encompass, which discrete mixing sites should be distinguished, and which ions should be considered to mix on each site; for example, Table A1. These choices determine which end-members will be required.
2. The $G(P, T)$ relations for those model end-members that are found in the appropriate version of the Holland and Powell (2011) internally consistent data set of thermodynamic properties of end-members (IDE).
3. $G(P,T)$ relations for those model end-members that are not in the IDE. For non-IDE end-member i, this is constructed from the $G(P, T)$ curves of a subset of Λ IDE end-members, as $G_i(P,T) = \Sigma_{\lambda=1}^{\Lambda} \nu_\lambda G_\lambda(P,T) + a + bT + cP$, where a, b, and c are constants and the net composition of the combination of end-members Λ yields the composition of i. The IDE end-members Λ do not necessarily appear anywhere else in the thermodynamic data set.
4. Activity-composition (a–x) relations, which describe the thermodynamic contributions of mixing among the end-members. In general, these follow the asymmetric formalism of Holland and Powell (2003). In the asymmetric formalism, the configurational entropy is formulated in terms of mixing on sites, potentially with a nonunity scaling factor applied to each site as described below. The nonideal mixing contribution from each end-member is defined in terms of a single interaction energy (Margules parameter, W) between each pair of end-members (Equation A20), which may be asymmetric and may be linearly dependent on P and/or T (thereby potentially contributing excess volume and/or entropy terms in addition to excess enthalpy).

In the Holland et al. (2018) model for the olivine solution, the end-members and mixing site occupancies are as shown in Table A1.

The end-members forsterite, fayalite, and monticellite are taken directly from the IDE. End-member cfm represents full ordering of Mg and Fe on the M1 and M2 sites; it is not in the IDE but the $G_{\mathrm{cfm}}(P, T)$ curve is formed from

$$G_{\mathrm{cfm}}(P,T) = 1/2\,(G_{\mathrm{fo}}(P,T) + G_{\mathrm{fa}}(P,T)) + \Delta G^{od}_{\mathrm{cfm}} \tag{A1}$$

where $\Delta G^{od}_{\mathrm{cfm}}$ is the Gibbs energy of ordering in the cfm composition, and has the form $a + bT + cP$.

Compositional and order variability within the solid solution are defined in terms of site fractions:

$$x = (x\mathrm{FeM1} + x\mathrm{FeM2})/(x\mathrm{FeM1} + x\mathrm{FeM2} + x\mathrm{MgM1} + x\mathrm{MgM2}) \tag{A2}$$

Table A1
End-Members and Mixing Site Occupancies of Olivine in the Holland et al. (2018) Thermodynamic Data Set

			Mixing sites				
			M1		M2		
End-member	Abbreviation	Formula	Mg	Fe	Mg	Fe	Ca
Forsterite	fo	Mg_2SiO_4	1	0	1	0	0
Fayalite	fa	Fe_2SiO_4	0	1	0	1	0
Ordered intermediate	cfm	$MgFeSiO_4$	1	0	0	1	0
Monticellite	mont	$CaMgSiO_4$	1	0	0	0	1

$$c = x\text{CaM2} \tag{A3}$$

$$Q = x - x\text{FeM1}/(x\text{FeM1} + x\text{MgM1}) \tag{A4}$$

where, for example, xFeM1 is the fraction of Fe on the M1 site. Compositional and order variables are subject to bounds, which for x, c, and Q, as in most cases, are [0.0 1.0]. The variables are chosen so as to ensure that the fraction of mixing ions on each site are normalized to a constant total and, if relevant, that charge balance is obeyed within the mineral.

Written in these variables, the site fractions are expressed as

$$x\text{MgM1} = 1 + Q - x, \tag{A5}$$

$$x\text{FeM1} = -Q + x, \tag{A6}$$

$$x\text{MgM2} = 1 - c - Q - x + cx, \tag{A7}$$

$$x\text{FeM2} = Q + x + (-c)x, \tag{A8}$$

$$x\text{CaM2} = c, \tag{A9}$$

The site fraction expressions are required to express the ideal activity, and hence the configurational entropy, of the model end-members in the solution. For some minerals, though not in olivine, the entropic contribution of a particular site is reduced by a scaling factor (see e.g., Holland et al., 2021) to simulate the effects of short-range order in the crystal structure. Thus, the ideal activity of end-member i can be expressed as

$$a^{id}_{i(\lambda)} = c_i \prod_s \left(X^s_{e_{s,i}}\right)^{\nu_s f_s} \tag{A10}$$

where $X^s_{e_{s,i}}$ is the site fraction of the element $e_{s,i}$ that appears on site s, ν_s is the number of atoms mixing on s, c_i the normalization constant to give $a^{id}_{i(\lambda)} = 1$ for pure i, and f the scaling factor for the site. For olivine ($f = 1$ for all sites), the ideal activities are

$$a^{id}_{\text{mont}} = x\text{MgM1} \cdot x\text{CaM2}, \tag{A11}$$

$$a^{id}_{\text{fa}} = x\text{FeM1} \cdot x\text{FeM2}, \tag{A12}$$

$$a^{id}_{\text{fo}} = x\text{MgM1} \cdot x\text{MgM2}, \tag{A13}$$

$$a^{id}_{\text{cfm}} = x\text{MgM1} \cdot x\text{FeM2}. \tag{A14}$$

The proportions of the end-members present at a given composition and state of order can be expressed as:

$$p_{\text{mont}} = c, \tag{A15}$$

$$p_{\text{fa}} = -Q + x, \tag{A16}$$

$$p_{\text{fo}} = 1 - c - Q - x + cx, \tag{A17}$$

$$p_{\text{cfm}} = 2Q + (-c)x. \tag{A18}$$

The equation of state is conveniently assembled in terms of the $\mu_i(P, T, \text{x}, \mathbf{Q})$, where μ_i is the chemical potential of end-member i, and $\mu_i = \frac{\partial G_i}{\partial p_i}$. μ_i can be written as

$$\mu_i = \mu^0_i + RT \log\left(a^{id}_i\right) + \mu^{ex}_i, \tag{A19}$$

where μ^0_i is the chemical potential of pure i, with $\mu^0_i(P, T) = G_i(P, T)$, as described above. In the asymmetric formalism, μ^{ex}_i is given by

$$\mu^{ex}_i = -\sum_{m=1}^{N_{ol}-1} \sum_{n>m}^{N_{ol}} \left(\phi'_m - \phi_m\right)\left(\phi'_n - \phi_n\right) W_{m,n} \left(\frac{2v_i}{v_m + v_n}\right). \tag{A20}$$

Table A2
Values of Parameters in the Model Olivine Solid Solution of Holland et al. (2018)

$W_{m,n}$ binary	Value
W(mont,fa)	24 kJ
W(mont,fo)	38 kJ
W(mont,cfm)	24 kJ
W(fa,fo)	9 kJ
W(fa,cfm)	4.5 kJ
W(fo,cfm)	4.5 kJ
v_{fo}	1
v_{fa}	1
v_{cfm}	1
v_{mont}	1
$\Delta G^{od,a}_{\mathrm{cfm}}$	0 kJ
$\Delta G^{od,b}_{\mathrm{cfm}}$	0 kJ/K
$\Delta G^{od,c}_{\mathrm{cfm}}$	0 kJ/kbar

Here, ϕ_i is the proportion of end-member i weighted by the asymmetry parameters, as $\phi_i = (p_i v_i) \Big/ \left(\sum_{m=1}^{N_{ol}} p_m v_m\right)$, with v_i the asymmetry parameter for end-member i. ϕ'_m is the value of ϕ_m in end-member i, such that $\phi'_m = 1$ where $m = i$ and $\phi'_m = 0$ where $m \neq i$. $W_{m,n}$ is the interaction energy between end-members m and n in the solution. The values of model parameters in the olivine solid solution are given in Table A2.

Appendix B: Implementation in MAGEMin

Mineral Assemblage Gibbs Energy Minimizer uses the input outlined in Appendix A to assemble the Gibbs energy of olivine at pressure P and temperature T:

$$G_{ol}(\mathbf{x},\mathbf{Q})|_{P,T} = \sum_{i=1}^{N_{ol}} \left(\mu_{i(ol)}(\mathbf{x},\mathbf{Q}) \cdot p_{i(ol)}(\mathbf{x},\mathbf{Q})\right)|_{P,T}, \tag{B1}$$

where $\mu_{i(ol)}$ is obtained as in Equation A20. Equation B1 constitutes the objective function for the minimization step. The gradient of the objective function is the derivative of the Gibbs from energy of olivine with respect to the compositional variables:

$$\frac{\partial G_{ol}}{\partial x_k} = \sum_{i=1}^{N_{ol}} \mu_{i(ol)} \frac{\partial p_{i(ol)}}{\partial x_k}, \tag{B2}$$

where $\frac{\partial p_{i(ol)}}{\partial x_k}$ is given in Table B1.

Table B1
Partial Derivatives of End-Member Proportions as a Function of Compositional Variables

	$\frac{\partial p_i}{\partial x}$	$\frac{\partial p_i}{\partial c}$	$\frac{\partial p_i}{\partial Q}$
$\frac{\partial p_{fo}}{\partial x_k}$	$c-1$	$x-1$	-1
$\frac{\partial p_{fa}}{\partial x_k}$	1	0	-1
$\frac{\partial p_{cfm}}{\partial x_k}$	$-c$	$-x$	2
$\frac{\partial p_{mont}}{\partial x_k}$	0	1	0

Table B2
Partial Derivatives of Site-Fractions as a Function of Compositional Variables

	$\frac{\partial sf_i}{\partial x}$	$\frac{\partial sf_i}{\partial c}$	$\frac{\partial sf_i}{\partial Q}$
$\frac{\partial xMgM1}{\partial x_k}$	-1	0	1
$\frac{\partial xFeM1}{\partial x_k}$	1	0	-1
$\frac{\partial xMgM2}{\partial x_k}$	$c-1$	$x-1$	-1
$\frac{\partial xFeM2}{\partial x_k}$	$1-c$	x	1
$\frac{\partial xCaM2}{\partial x_k}$	0	1	0

During the minimization, the value of all site fractions is required to be ≥0 via a set of nonlinear inequality constraints (derived from Equation A5 to A9) that is passed to the local minimizer. The gradients of the site-fractions with respect to the compositional and order variables are given in Table B2.

The above expressions are passed to NLopt (Johnson, 2021) together with an initial guess for the compositional variables. Subsequently, the objective function is minimized using the CCSAQ algorithm (Svanberg, 2002).

Data Availability Statement

A complete guide on how to download, install and run MAGEMin is given in the Git repository https://github.com/ComputationalThermodynamics/magemin.git. The version of the code used to produce the pseudosections is also made available on Zenodo, https://doi.org/10.5281/zenodo.6347567.

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